



Student Number	
Mark / 48	

Chemistry

2008 Half Yearly Year 11
Examination

Theory and
Data Processing

General Instructions

- Reading time – 5 minutes
 - Working time – 80 minutes
 - Write using black or blue pen
 - Write your Student Number at the top of this page and on the response sheet on page 7.
- A data sheet and a periodic table are provided.

Theory

Total Marks – 48

Part A – 12 marks

Attempt Questions 1 – 12

Part B – 37 marks

Attempt Questions 13-21

Data Processing

Total Marks - 7

Part A- Multiple Choice 12 marks
Attempt Questions 1-12
Allow about 15 minutes for this part

Select the alternative A, B, C or D that best answers the question. Fill in the response oval completely.

Sample: $2 + 4 =$ (A) 2 (B) 6 (C) 8 (D) 9
A B C D

If you think you have made a mistake, put a cross through the incorrect answer and fill in the new answer.

A B C D

If you change your mind and have crossed out what you consider to be the correct answer, then indicate the correct answer by writing the word *correct* and drawing an arrow as follows.

A B C D
correct →

▶ Mark your answers for Questions 1- 12 in the Answer Box on page 7

1. Which of the following lists the Ages in history from earliest to latest?
- (A) Stone Age, Copper Age, Brass Age, Iron Age.
 - (B) Stone Age, Bronze Age, Copper Age, Iron Age.
 - (C) Stone Age, Copper Age, Bronze Age, Iron Age.
 - (D) Iron age, Copper Age, Bronze Age, Stone Age.
2. Sydney's main water supply comes from Warragamba dam. Before being distributed to Sydney households, the water is first purified in one of the various purification plants in the outskirts of the city. What process do you think is being used to remove relatively large (pebble sized) debris from the water?
- (A) sieving
 - (B) evaporation
 - (C) distillation
 - (D) filtration
3. Given the following processes
- (i) production of oxygen gas when mercury (II) oxide is heated.
 - (ii) production of water vapour when water is boiled.
 - (iii) production of misty cloud when dry ice is placed in water.
 - (iv) production of water droplets on the outside of a soft drink bottle when taken out of the fridge.

Which of the processes is a chemical change?

- (A) (i) only
- (B) (i) & (iii) only
- (C) (i), (iii) & (iv) only
- (D) (i), (ii), (iii) & (iv)

4. Which of the following substances has the greatest ionic bond strength?

	<i>Substance</i>	<i>Melting Point (°C)</i>
(A)	ZnO	1970
(B)	HgO	d* 500
(C)	MgO	2800
(D)	graphite	3730

* decomposes

5. Which of the following is not a molecule?

- (A) He
- (B) N₂
- (C) NH₃
- (D) Al

6. Of the four ions below, which of the following statements is true?



- (A) The number of neutrons is the same in each.
- (B) They are all members of the same periodic group.
- (C) They all belong to the same period.
- (D) They all contain the same number of electrons.

7. The characteristics of a certain solid, X, is detailed in the table below along with some other known substances.

<i>Substance</i>	<i>Appearance</i>	<i>Conductivity ($MS\ m^{-1}$) at $25^{\circ}C$</i>
carbon (graphite)	dull black	0.07
magnesium	silvery	23
phosphorus	dull white	10^{-15}
X	silvery-white	10^{-3}

What is X likely to be?

- (A) metal
- (B) non-metal
- (C) semi-metal
- (D) noble gas

8. Which of the following best describes the use of solder in terms of its properties?

	Use	Property
(A)	Door knobs	Resists corrosion
(B)	Joining wires	Low melting point
(C)	Church bells	Easily cast
(D)	Fillings for teeth	Easily worked

9. The physical properties of solid metals can best be explained by which of the following?
- (A) Each atom is bonded in the crystal lattice by covalent bonds.
 - (B) Positive metal ions are arranged in an orderly way, with valence electrons able to move freely through the crystal lattice.
 - (C) Positive and negative metal ions are arranged in an orderly way, with mobile valence electrons able to migrate easily around the crystal lattice.
 - (D) Each metal atom is surrounded by a variable number of valence electrons which complete the "noble gas" electronic structure in the crystal lattice.
10. Sodium metagermanate has the formula Na_2GeO_3 . What is the chemical formula of indium metagermanate?
- (A) InGeO_3
 - (B) $\text{In}_2(\text{GeO}_3)_3$
 - (C) In_3GeO_3
 - (D) $\text{In}(\text{GeO}_3)_3$
11. Which of the following exists as a covalent molecular substance in the solid state?
- (A) bromine
 - (B) sodium chloride
 - (C) diamond
 - (D) aluminium

12. Refer to the modified periodic table below to answer this question

I	II		III	IV	V	VI	VII	Noble Gases
				P			U	
	Q		R		S	T		
								X
Y								

Which of the following pairs would form the most ionic compound?

- (A) Q and X
- (B) R and P
- (C) S and T
- (D) Y and U

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Part A: Answer grid for multiple choice questions.

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|-----|-----|-----|-----|-----|
| 1. | A O | B O | C O | D O |
| 2. | A O | B O | C O | D O |
| 3. | A O | B O | C O | D O |
| 4. | A O | B O | C O | D O |
| 5. | A O | B O | C O | D O |
| 6. | A O | B O | C O | D O |
| 7. | A O | B O | C O | D O |
| 8. | A O | B O | C O | D O |
| 9. | A O | B O | C O | D O |
| 10. | A O | B O | C O | D O |
| 11. | A O | B O | C O | D O |
| 12. | A O | B O | C O | D O |

Mark	
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Part B. 36 marks

Attempt Questions 13 - 21

Allow about 55 minutes for this part

▶ Show all relevant working in questions involving calculations.

Question 13 (4 marks)

A particularly keen JRAHS student obtained water samples from the *Dead Sea*, while holidaying in Jerusalem. He wanted to compare the percentage total salt of this seawater with that of his hometown's. Using school equipment he determined the percentage of total salt in his *Dead Sea* water sample. The results of his analysis is detailed below:

mass of evaporating basin = 253.3 g
mass of evaporating basin + seawater = 766.2 g
mass of salt residue = 165.0 g

- (a) Calculate the percentage composition of the seawater sample. Show your working. (3 marks)

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- (b) State one possible source of error for this determination. (1 mark)

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Question 14 (4 marks)

Describe one use of copper metal and one of helium and justify their use(s) on the basis of their physical properties.

(a) copper (2 marks)

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(b) helium (2 marks)

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Question 15 (5 marks)

Magnesium is a common metal used in flash bulb photography.

(a) Explain the valency and reactivity of magnesium in terms of its electronic configuration. (2 marks)

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(b) Use Lewis electron dot structures to illustrate the reaction of magnesium with the oxygen atom. (2 marks)

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(c) Draw the Lewis electron dot structure for the oxygen molecule. (1 mark)

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Question 16 (3 marks)

Complete the table to illustrate the contrast between boiling and electrolysis of water in terms of physical & chemical change, breaking of a chemical bond or intermolecular forces and the relative amount of the energy required.

	<i>Boiling</i>	<i>Electrolysis</i>
Type of change		
Type of forces broken		
Magnitude of energy involved in the process (high/low)		

Question 17 (4 marks)

A variety of chemical equations can be used to describe chemical reactions.

- (a) Give a word equation for the chemical reaction that occurs when sulfur is strongly heated in oxygen. (1 mark)

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- (b) Give a balanced formulae equation for the reaction of sodium in water. (1 mark)

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- (c) Give the half equations for the reaction of iron metal with hydrochloric acid. (2 marks)

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Question 18 (3 marks)

Metals have been extracted from their ores and used for many thousands of years.

- (a) Give the formula for an ore from which iron is extracted. (1 mark)

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- (b) Explain why energy is necessary to extract a metal from its ore. (2 marks)

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Question 19 (3 marks)

When Mendeleev published his Periodic Table, he predicted the existence of three elements not yet discovered, Ekaaluminium, Ekasilicon, and Ekaboron.

These elements, later identified as Gallium, Germanium and Scandium respectively, were subsequently discovered and found to have properties very similar to those he had predicted.

- (a) What formulae do you think Mendeleev may have predicted for the oxides of Germanium and Gallium? (1 mark)

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- (b) Outline on what basis Mendeleev made these predictions. (2 marks)

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Question 20 (5 marks)

Refer to the table below to answer question 20.

Substance	MP (°C)	Electrical conductivity	
		Solid state	Liquid state
Q	-102	Nil	Nil
X	1423	Nil	Conducts
Y	3600	Nil	Nil
Z	1535	Conducts	Conducts

- (a) Which substance is iron?
- (b) Which substance is calcium fluoride?.....
- (c) Which substance is dinitrogen trioxide?.....
- (d) Which substance is diamond?
- (e) For substance X, identify the type of particle that carries the current in the liquid state.....
- (f) For substance Z, identify the type of particle that carries the current in the liquid state.....

Question 21 (5 marks)

Chlorine is a reactive element in Group VII of the periodic table. Chlorine forms compounds with nearly all of the other elements. With oxygen, several such compounds are known, one of which is the heptoxide Cl_2O_7 .

- (a) Identify the trend in the reactivity of the elements in Group VII. (1 mark)

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- (b) Explain the overall trend in the ionization energy of the elements in the period ending with chlorine. (3 marks)

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- (c) Using your knowledge of valencies and the Periodic Table, predict the formula of another compound (other than Cl_2O_7) formed between chlorine and oxygen. (1 mark)

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End Of Test